Key points to learn

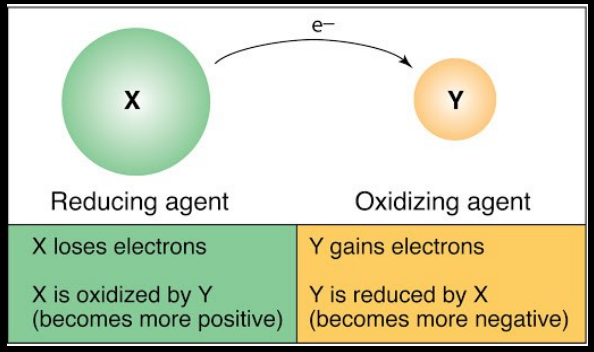
1. Oxidation & reduction	Oxidation Is Lost of electrons Reduction Is Gain of electrons
2. Rules for oxidation state	-Work out the oxidation state of the know elements -Work out the oxidation state of the unknow elements.
3. The know oxidation state	H= +1 O= -1
4. Redox	Oxidation and reduction reactions are reactions in which one species is reduced and another is oxidized at the same time. Both side should be cancelling
5. Disproportion reaction	When one species in a reaction undergoes both oxidation and reduction.

Key points to learn

	2Fe2O3 (s) + 3C (s) —> 4Fe (s)
6. Working out redox reaction	Look for know element oxidation. Single element have state 0. Then look for unknow element oxidation state. See if the element have been oxidised or reduced.
7. Oxidation &reduction term of oxygen transfer	Oxidation is gain of oxygen Reduction is lost of oxygen
8. Oxidising agent	Gain of electrons/themselves reduces
9. Reducing agent	Lost of electrons/ themselves oxidised
10. Step to balance simple redox reaction	Check both side has equal charges. Balance the number of electrons. Deduce if the reaction is oxidised or reduced
11. Harder redox half equation	1.Balance the number of atoms on each side. 2.number of oxygen by adding H20. 3.no. of hydrogen by adding H+. 4.charges on each side of the equation by adding electrons.

Elements of the Sea

Knowledge organiser



Background information

Did you know: Glass is actually a liquid, it just flows very, very slowly

Additional information

Oscar **A**te **G**reen **O**ranges = **O**xidising **A**gents **G**ive **O**xygen

Ruby **A**te **R**ed **O**ranges = **R**educing **A**gainst **R**emove **O**xygen